

Time: 3 hours

Max. Marks: 80

- N.B:
1. Question.No.1 is compulsory.
 2. Attempt any three questions from Q.No.2 to Q.No.6
 3. Assume suitable data.
 4. Figures to the right indicate full marks

Q1 Solve any Four out of Six (5 Marks each) (20 Marks)

- A Write a short note on Electroosmosis.
- B Describe acid base catalysis in detail.
- C Write short note on Differential Thermal Analysis (DTA).
- D Write any 2 preparations and properties of Malonic ester.
- E Explain the importance of non-aqueous solvents with 2 examples.
- F Explain GC-MS spectroscopy.

Q2 (20 Marks)

- A Give the basic requirement of IR radiation absorption. Give any two applications of IR spectroscopy.
- B Explain Paper chromatography in detail.
- C Define non aqueous solvents. What are advantages & disadvantages of liquid ammonia as a solvent?
- D Give the Mechanism of Reformatsky Reaction, with one example.

Q3 (20 Marks)

- A Write Intermediate compound formation theory for catalysts.
- B What are Auxochrome & chromophores? Give 2 applications of UV spectroscopy.
- C Explain emulsions & their types.
- D What do you mean by aromatic character? Explain it for Pyridine with any 2 reactions.

Q4 (20 Marks)

- A Give a note on Donnan membrane equilibrium & its significance.
- B Explain in detail column chromatography.
- C What are surfactants? Explain their role as a colloids.
- D Explain Acid base reactions & Redox reactions in case of Liq. SO₂.

Q5 (20 Marks)

- A Explain concept of electrical double layer with Helmholtz & Stern model.
- B Write one preparation & two applications of Acetoacetic ester.
- C Write Principle of NMR spectroscopy. Explain Shielding & Deshielding of protons.
- D Write a note on Gas solid chromatography.

Q6 (20 Marks)

- A What are levelling effects of non-aqueous solvents?
- B Write in detail how to determine purity of given sample using TLC technique
- C What are colloids? Give its importance in Foods.
- D Explain action of promoters & catalytic poisons.
